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# 4-Amino-5-(2-hydroxy-4,4-dimethyl-6-oxocyclohexenylmethyl)-1-methyl-2-(methylsulfanyl)pyrimidin-6(1H)one: hydrogen-bonded chains are $\pi$-stacked in pairs 

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Molecules of the title compound, $\mathrm{C}_{15} \mathrm{H}_{21} \mathrm{~N}_{3} \mathrm{O}_{3} \mathrm{~S}$, have a markedly polarized electronic structure; the carbocyclic ring adopts an envelope conformation and the overall molecular conformation appears to be controlled by two intramolecular hydrogen bonds, one each of the $\mathrm{O}-\mathrm{H} \cdots \mathrm{O}$ and $\mathrm{N}-\mathrm{H} \cdots \mathrm{O}$ types. The molecules are linked into $C(6)$ chains by an intermolecular $\mathrm{N}-\mathrm{H} \cdots \mathrm{O}$ hydrogen bond, and pairs of these hydrogen-bonded chains are linked by a $\pi-\pi$ stacking interaction.

## Comment

As part of a program for the synthesis of fused pyrimido derivatives, we have investigated the three-component cyclocondensation between 4-amino-1-methyl-2-(methylsulfanyl)-pyrimidin-6(1H)-one, formaldehyde and 5,5-dimethylcyclo-hexane-1,3-dione (dimedone), but instead of the desired pyrimido $[3,4-b] q u i n o l i n e$, the intermediate title compound, (I), was obtained. We report here its molecular and supramolecular structure. Three-component reactions of this general type can also provide, as a by-product, the twofold symmetrical compound $2,2^{\prime}$-methylenebis(5,5-dimethyl-3-hydroxy-2-cyclohexen-1-one), (II), whose structure we have reported recently (Low et al., 2003).

Within the heterocyclic ring of (I), the C4-C5 and C5-C6 distances (Table 1) are similar, and they cannot be associated readily with double and single bonds, respectively. In addition, the $\mathrm{C} 4-\mathrm{N} 4$ distance is very much shorter than the mean value

[^0](1.355 $\AA$; Allen et al., 1987) for bonds of type C(aryl) $-\mathrm{NH}_{2}$, where N is planar. In contrast to the $\mathrm{C}-\mathrm{C}$ bonds, the $\mathrm{C} 2-\mathrm{N} 3$ and N3-C4 bonds have distances typical of those for localized double and single bonds of these types. While the C6O6 distance is typical of those in unstrained amides, the C6N 1 distance is very long for its type (mean value $1.346 \AA$; Allen et al., 1987). These values, taken together, point to the significance of the polar form ( $\mathrm{I} a$ ) as an important contributor to the overall molecular-electronic structure but effectively eliminate as contributors other forms, such as (Ib) and (Ic).

(I)

(II)

(Ia)

( $1 b$ )

(Ic)

There is clear evidence for bond fixation in the atom sequence between O52 and O56. Within the carbocyclic ring, the total puckering amplitude, $Q$, is 0.455 (2) $\AA$, markedly less than the value $(0.63 \AA)$ for an ideal cyclohexane chair of $\overline{3} m$ $\left(D_{3 d}\right)$ molecular symmetry, with $\mathrm{C}-\mathrm{C}$ distances of $1.54 \AA$ (Cremer \& Pople, 1975). For the atom sequence C51 to C56, the angular parameters $\left[\theta=53.9(3)^{\circ}\right.$ and $\varphi=195.8$ (4) $\left.{ }^{\circ}\right]$ are clearly indicative of the sofa or envelope conformation being the best single qualitative descriptor, with a local pseudomirror plane passing through atoms C5, C54, C541 and C542 (Fig. 1); this conformation is confirmed by the leading torsion angles involving this ring (Table 1). However, the envelope conformation can itself be represented as a linear combination of the three primitive forms chair, boat and twist-boat (Evans \& Boeyens, 1989), with the chair form dominant here.

There are three hydrogen bonds (Table 2), two of which are intramolecular, and these doubtless control the overall molecular conformation, in particular, the twist angles about the C5-C57 and C51-C57 bonds (Table 1). The short O $\cdots \mathrm{O}$ distance in the intramolecular $\mathrm{O}-\mathrm{H} \cdots \mathrm{O}$ hydrogen bond is consistent with the development of a significant negative charge on atom $\mathrm{O} 6[c f$. form ( $\mathrm{I} a)]$. Each of the intramolecular hydrogen bonds gives rise to an independent $S(8)$ motif (Bernstein et al., 1995). The molecules are linked into chains by a single intermolecular $\mathrm{N}-\mathrm{H} \cdots \mathrm{O}$ hydrogen bond; amine atom N 4 in the molecule at $(x, y, z)$ acts as a hydrogen-bond donor, via atom $\mathrm{H} 4 B$, to amide atom O6 in the molecule at


Figure 1
The molecule of (I), showing the atom-labelling scheme. Displacement ellipsoids are drawn at the $30 \%$ probability level.


Figure 2
Part of the crystal structure of (I), showing the formation of a $C(6)$ chain along [010]. For clarity, H atoms bonded to C atoms have been omitted. Atoms marked with an asterisk (*) or a hash (\#) are at the symmetry positions $(x, 1+y, z)$ and $(x,-1+y, z)$, respectively.


Figure 3
A stereoview of part of the crystal structure of (I), showing the $\pi$-stacked double chain in the domain $0.24<z<0.76$. For clarity, H atoms bonded to C atoms have been omitted.
$(x, 1+y, z)$, so generating by translation a $C(6)$ chain running parallel to the [010] direction (Fig. 2).

Four of these chains pass through each unit cell, and they are linked into pairs by a single $\pi-\pi$ stacking interaction. The pyrimidine rings of the molecules at $(x, y, z)$ and $(1-x, 1-y$, $1-z$ ) are parallel, with an interplanar spacing of 3.498 (2) $\AA$. The centroid separation is 3.646 (2) $\AA$, corresponding to a centroid offset of 1.028 (2) $\AA$. The resulting pair of $\pi$-stacked chains (Fig. 3) lies in the domain $0.24<z<0.76$, and a second double chain, related to the first by the translational symmetry operators, lies in the domain $0.74<z<1.26$. There are no direction-specific interactions between the double chains.

## Experimental

To a hot solution of 4-amino-1-methyl-2-(methylsulfanyl)pyrimidin$6(1 H)$-one ( 2 mmol ) in ethanol ( $20 \mathrm{ml}, 96 \%$ ) were added formaldehyde ( 2 mmol ) and dimedone ( 2 mmol ); the reaction mixture was heated under reflux for 4 h . The white powder that formed during the reaction was filtered off and the remaining solution was cooled to 270 K. After 2 d, a white solid was filtered off and recrystallized from ethanol, affording crystalline needles suitable for single-crystal X-ray diffraction (yield $30 \%$; m.p. 426-428 K). Analysis found: C 55.7, H 6.7, N 12.8, S 9.9\%; $\mathrm{C}_{15} \mathrm{H}_{21} \mathrm{~N}_{3} \mathrm{O}_{3} \mathrm{~S}$ requires: C 55.7, H 6.6, N 13.0, S $9.9 \%$.

## Crystal data

$\mathrm{C}_{15} \mathrm{H}_{21} \mathrm{~N}_{3} \mathrm{O}_{3} \mathrm{~S}$
$M_{r}=323.41$
Monoclinic, $P 2_{1} / n$
$a=7.5524$ (2) $\AA$
$b=7.1370(3) \AA$
$c=28.6208$ (11) $\AA$
$\beta=93.128$ (2) ${ }^{\circ}$ 。
$V=1540.4$ (1) $\AA^{3}$
$Z=4$
$D_{x}=1.394 \mathrm{Mg} \mathrm{m}^{-3}$
Mo $K \alpha$ radiation
Mo $K \alpha$ radiation
Cell parameters from 3586
reflections
$\theta=3.4-28.2^{\circ}$
$\mu=0.23 \mathrm{~mm}^{-1}$
$T=120$ (2) K
Needle, colourless
$0.50 \times 0.30 \times 0.20 \mathrm{~mm}$

## Data collection

Nonius KappaCCD diffractometer $\varphi$ scans, and $\omega$ scans with $\kappa$ offsets Absorption correction: multi-scan
(SORTAV; Blessing, 1995, 1997)
$T_{\text {min }}=0.868, T_{\text {max }}=0.956$
18675 measured reflections 3586 independent reflections

$$
\begin{aligned}
& 2600 \text { reflections with } I>2 \sigma(I) \\
& R_{\text {int }}=0.071 \\
& \theta_{\max }=28.2^{\circ} \\
& h=-9 \rightarrow 9 \\
& k=-8 \rightarrow 9 \\
& l=-37 \rightarrow 37
\end{aligned}
$$

## Refinement

Refinement on $F^{2}$
$R\left[F^{2}>2 \sigma\left(F^{2}\right)\right]=0.042$
$w R\left(F^{2}\right)=0.110$
$S=1.01$
3586 reflections
204 parameters
H -atom parameters constrained

$$
\begin{aligned}
& w=1 /[ \sigma^{2}\left(F_{o}^{2}\right)+(0.0505 P)^{2} \\
&+0.5167 P] \\
& \text { where } P=\left(F_{o}^{2}+2 F_{c}^{2}\right) / 3 \\
&(\Delta / \sigma)_{\max }=0.001 \\
& \Delta \rho_{\max }=0.28 \mathrm{e} \AA^{-3} \\
& \Delta \rho_{\min }=-0.29 \mathrm{e}^{-3}
\end{aligned}
$$

Table 1
Selected geometric parameters ( $\left(\AA{ }^{\circ}\right)$.

| N1-C2 | $1.334(2)$ | N1-C11 | $1.447(2)$ |
| :--- | :---: | :--- | ---: |
| C2-N3 | $1.282(2)$ | C4-N4 | $1.319(2)$ |
| N3-C4 | $1.375(2)$ | C6-O6 | $1.234(2)$ |
| C4-C5 | $1.359(3)$ | C51-C52 | $1.343(3)$ |
| C5-C6 | $1.378(3)$ | C52-O52 | $1.317(2)$ |
| C6-N1 | $1.404(2)$ | C51-C56 | $1.423(3)$ |
| C2-S2 | $1.754(2)$ | C56-O56 | $1.227(2)$ |
| S2-C21 | $1.750(2)$ |  |  |
|  |  |  | $72.9(2)$ |
| C4-C5-C57-C51 | $78.6(2)$ | C5-C57-C51-C52 | $5.7(2)$ |
| C56-C51-C52-C53 | $3.4(3)$ | C52-C51-C56-C55 | $-36.1(2)$ |
| C51-C52-C53-C54 | $18.6(3)$ | C54-C55-C56-C51 | $-34.6(2)$ |
| C52-C53-C54-C55 | $-45.3(2)$ | C53-C54-C55-C56 | 54.6 |
| C52-C53-C54-C541 | $74.5(2)$ | C56-C55-C54-C541 | $-65.4(2)$ |
| C52-C53-C54-C542 | $-164.09(15)$ | C56-C55-C54-C542 | $173.8(2)$ |
|  |  |  |  |

Table 2
Hydrogen-bonding geometry $\left(\AA,{ }^{\circ}\right)$.

| $D-\mathrm{H} \cdots A$ | $D-\mathrm{H}$ | $\mathrm{H} \cdots A$ | $D \cdots A$ | $D-\mathrm{H} \cdots A$ |
| :--- | :--- | :--- | :--- | :--- |
| O52-H52 $\cdots$ O6 | 0.84 | 1.87 | $2.681(2)$ | 161 |
| N4-H4A O56 | 0.88 | 2.13 | $2.952(2)$ | 155 |
| N4-H4B $\cdots$ O6 $^{\mathrm{i}}$ | 0.88 | 2.14 | $2.831(3)$ | 134 |

Symmetry code: (i) $x, 1+y, z$.

Space group $P 2_{1} / n$ was uniquely assigned from the systematic absences. All H atoms were located from difference maps and were thereafter treated as riding atoms, with $\mathrm{C}-\mathrm{H}$ distances of $0.98\left(\mathrm{CH}_{3}\right)$ and $0.99 \AA\left(\mathrm{CH}_{2}\right), \mathrm{N}-\mathrm{H}$ distances of $0.88 \AA$ and an $\mathrm{O}-\mathrm{H}$ distance of 0.84 Å.

Data collection: KappaCCD Server Software (Nonius, 1997); cell refinement: DENZO-SMN (Otwinowski \& Minor, 1997); data reduction: DENZO-SMN; program(s) used to solve structure: OSCAIL (McArdle, 2003) and SHELXS97 (Sheldrick, 1997); program(s) used to refine structure: SHELXL97 (Sheldrick, 1997); molecular graphics: PLATON (Spek, 2003); software used to prepare material for publication: SHELXL97 and PRPKAPPA (Ferguson, 1999).

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Supplementary data for this paper are available from the IUCr electronic archives (Reference: SK1697). Services for accessing these data are described at the back of the journal.

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